1) Briefly describe the primary ideas behind VSEPR theory.

- 2) For each of the following compounds, a Lewis structure, determine the bond angles and molecular shapes for all atoms:
 - a) BI_3

b) CH₄

c) NF_3

d) C₂H₂

- 1) Electrons are negatively charged and like repels like so the atoms in a molecule will likely be as far away from each other as possible.
- 2) a) BI₃ 120° 120° ′120° trigonal planar b) CH_4 109.5° Ħ √ Η ′109.5° tetrahedral c) NF_3 107.5° -107.5° / 107.5°

trigonal pyramidal